***Lesson Plans for the Week of: 3/27/17 Teacher: Hough Course: Chemistry Period: 1,3,7/8***

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| Elements ofa Lesson | **Monday** | **Tuesday** | **Wednesday** | **Thursday** | **Friday** |
| Objective/Focus/Essential Question | CH.5bSpiral review of use of ICE diagramsUnderstand and apply Dalton’s Law of Partial Pressures | CH.5bUnderstand and utilize the Ideal Gas Law (IDG) | CH.5bCorrectly identify the correct gas law or method needed to solve these types of word problems, then solve | CH.2f,gSpiral Review: periodic table trends for atomic radius, ionization energy, and electronegativityUnderstand the following topics about the concept of shielding and its periodic table trend | CH.2e,g--Draw electron configuration any of the first 20 elements--understand the aufbau principle, Hund’s rule, and the Pauli Exclusion Principle |
| Lesson/Act.Type of Presentation | Weekly memorization (name and formula):CH4, C2H6, C2H4, C2H2, CH2O, C6H6Bellwork: Solve an ICE problem (a Boyle’s Law example) Go over another exampleWhole Group: (continuation from previous class period)There are 2 types of questions which utilize Dalton’s Law: model one of eachIndividual:Student Practice  | Bellwork: ICE example (using Charles’ Law)Whole group:Address questions about previous day’s workIntroduce ideal Gas Law—concept and equation, plus value for R and importance of units: volume in L; pressure in atm; and temp in KModel solving IGL problemIndividual:Student practice  | Whole group:Identify cues which can help indicate which method to use when solving gas law problemsIndividual:Students will solve a variety of gas law problems | Bellwork: ionization energy trend questionDiscussReview periodic table trends known so far (poorly on previous 2 tests)Whole group:Define electron shielding; and explain its periodic table trends Examples from the periodic table will be given to help illustrate the trends  | Whole group:Explain how to determine and describe the locations of electrons in an atom:a) explain how electrons are arranged in the atom (use a strange hotel with 4 floors and multiple subfloors as an example); Model how to determine the electron locations for a simple element like lithiumb) explain Aufbau principle, Hund’s Rule, and Pauli exclusion principlec) Explain how to write the electron configuration for a given elementd) Hand out spdf diagram and explain how it can be used as a mnemonice) Model samples: H, He, nitrogen, copperIndividual: e) Students write electron configuration for a few elements: beryllium, fluorine, sodium, chromium |
| Evaluation | Student work in class and on homework problems | Student work on practice problems | Accuracy of student work | Whole group formative assessment questions | Students write electron configurations for three elements |
| Extension/Homework | Practice word problems utilizing Dalton’s Law: worksheet | Practice word problems utilizing this principle: worksheet |  |  | Students write electron configurations: p. 136#8,9 |

MATERIALS:

Monday: Dalton’s Law worksheet

Tuesday: IDG worksheet

Wed.: Mixed gas laws worksheet

Thursday: periodic table trend worksheet

Friday: spdf diagram